

1. Bond breaking absorbs energy
AND bond making releases energy ✓
ALLOW bond breaking is endothermic
AND bond making is exothermic

More energy released than absorbed ✓

ALLOW exothermic change transfers more energy than endothermic change
OR bond making transfers more energy than bond breaking
OR 'the sum of the bond enthalpies in the products is greater than the (sum of the) bond enthalpies in the reactants'
OR 'the sum of the bond enthalpies of the bonds made is greater than (the sum of) the bond enthalpies of the bonds broken'
IGNORE reference to strong and weak bonds
IGNORE enthalpy of products is less than enthalpy of reactants

[2]

2. Respiration ✓

IGNORE anaerobic

[1]

3. (i) $100 \times 4.18 \times 17.3$ ✓

ALLOW 7231 J ✓

7.23 (kJ) ✓

ALLOW 7.23 with no working out
ALLOW from 7.2 up to calculator value of 7.2314
ALLOW from 0.060 up to calculator value for 1 mark
 (i.e. ECF from use of $m = 0.831$ in first stage)
IGNORE sign

2

- (ii) $M_r = 180$ ✓

amount = 4.62×10^{-3} (mol) ✓

ALLOW 4.6×10^{-3} *OR* 4.62×10^{-3}
OR 4.617×10^{-3} up to calculator value
DO NOT ALLOW 0.005
ALLOW ECF from wrong M_r

2

- (iii) $\Delta H_c = 1560 \text{ (kJ)}$ **OR** 1570 (kJ)
but answer must be to 3 sig fig ✓

*ALLOW ECF from 'answer to (i) $\div$ answer to (ii)'
but answer must be to 3 sig fig*

minus sign ✓

minus mark is an independent mark

2

[6]

4. +1250 ✓

ALLOW full marks for -2830 with no working out ✓✓✓

$+(-394 \times 6) + (-286 \times 6)$ **OR** -4080 ✓
 -2830 ✓

ALLOW for 2 marks:

+2830 cycle wrong way around

OR 1400 **OR** 860 one value not $\times 6$

OR -5330 **OR** $+5330$ wrong sign for 1250 or 4080

OR $+570$ ✓✓ correct cycle but not $\times 6$

ALLOW for 1 mark:

-1400 **OR** -860 cycle wrong way around and one value not $\times 6$

OR -570 cycle wrong way around and not $\times 6$

OR -1930 **OR** $+1930$ ✓ wrong sign and not $\times 6$

Note: There may be other possibilities.

[3]

5. Any two from the following:

Heat released to the surroundings ✓

ALLOW heat loss

Incomplete combustion **OR** incomplete reaction

OR not everything burns

IGNORE reference to evaporation

Non-standard conditions ✓

[2]

6. (i) Acid ✓

ALLOW correct formula if no name given:

e.g. H_3PO_4 **OR** H_2SO_4 **OR** H^+ ✓

ALLOW correct name of acid even if an incorrect formula is used

IGNORE heterogeneous **OR** homogeneous

1

(ii) The position of equilibrium will shift so as to minimise the effect of any change in conditions ✓

DO NOT ALLOW 'reaction shifts'

The idea of a shift in equilibrium is essential

1

(iii) Low temperature **AND** high pressure ✓

One mark for conditions.

This mark is independent of the reasons for conditions

Low temperature because the (forward) reaction is exothermic ✓

One mark for reason for the chosen temperature

High pressure because there are fewer moles (of gas) on the right hand side ✓

One mark for reason for the chosen pressure

ALLOW fewer moles of products

3

(iv) (60 atmosphere pressure is a) high pressure may be too expensive **OR** may cause safety problems ✓

(300 °C is sufficiently high) to give a fast rate of reaction ✓

without shifting equilibrium to the left

OR compromising equilibrium yield ✓

3

[8]

7. The enthalpy change for the complete combustion ✓

of 1 mol (of a substance) ✓

ALLOW energy change for combustion in excess oxygen **OR** energy released during complete combustion **OR** energy change for combustion in excess air

NOT energy required

This mark is not stand alone but must relate to statement about an enthalpy change even if the statement was not awarded a mark

[2]

8. (i) 56.430 (kJ) ✓

ALLOW 56.43 (kJ) **OR** 56.4 kJ ✓ **OR** 56 kJ
ALLOW -56.43 i.e. ignore sign

1

(ii) $M_r [\text{CH}_3(\text{CH}_2)_4\text{OH}] = 88.0$ ✓

$n = 0.0200$ mol ✓

ALLOW 88

ALLOW 0.02 **OR** ecf from wrong M_r

ALLOW full marks for 0.02 with no working out

2

(iii) (-2821.5) ✓

$= (-2820)$ (3 SF) ✓

correct minus sign ✓

ALLOW correct substitution into formula (b)(i) $\div$ (b)(ii) e.g.

$56.4 \div 0.02$ this is essentially a mark for the working

ALLOW ecf from i.e. answer from (b)(i) $\div$ (b)(ii)

The minus mark is stand alone and is independent of the numerical answer

3

[6]

9. (i) pressure: 100 kPa **OR** 101 kPa

AND

temperature: 298 K **OR** 25°C ✓

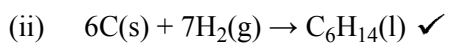
units needed

ALLOW 1 bar **OR** 1 atm **OR** 760 mmHg

ALLOW any stated temperature so for example 100kPa and 40°C would be credited with a mark

IGNORE any reference to moles or concentration

1



ALLOW graphite / gr

1

- (iii) many different hydrocarbons would form
OR activation energy too high
OR reaction too slow
OR they don't react together ✓
ALLOW can form different isomers OR can form different structures
IGNORE reaction may be reversible

1

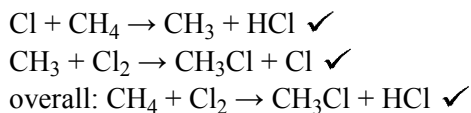
- (iv) $6 \times -394 + 7 \times -286$ shown **OR** calculated as -4366 ✓
 -4366 and -4163 added **OR** subtracted ✓
 correct answer $-4366 - (-4163) = -203$ ✓
ALLOW THREE marks for -203 on its own with no working out or written on the answer line
ALLOW TWO marks for $+203, +3483, +1513, +1767$ or -8529 on its own with no working out
ALLOW ONE mark for $-3483, -1513, -1767$ or $+8529$ on its own with no working out
units NOT needed
Positive sign not needed for endothermic answers

3

[6]

10. (i) $\text{Cl} + \text{O}_3 \rightarrow \text{ClO} + \text{O}_2$ ✓
 $\text{ClO} + \text{O} \rightarrow \text{Cl} + \text{O}_2$ ✓
overall: $\text{O}_3 + \text{O} \rightarrow 2\text{O}_2$ ✓

OR



Marks must come from one or other of the radical process and not from both of them.

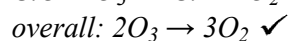
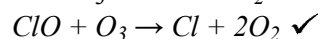
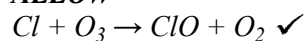
If two processes are described then an incorrect step in one process will contradict a correct step in the other process.

ALLOW overall equation mark even if the steps are wrong

*the radicals do **NOT** need a single dot*

IGNORE any state symbols

ALLOW



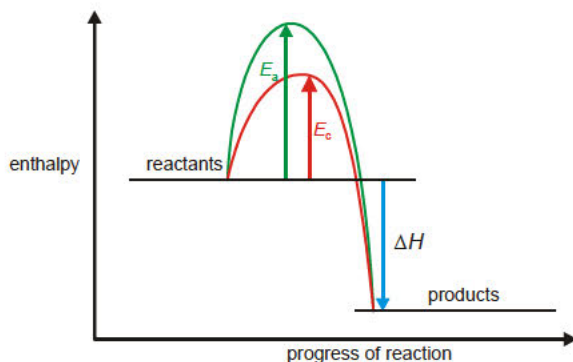
ALLOW any saturated hydrocarbon including cyclic

ALLOW ecf for second step and overall reaction if wrong hydrocarbon used e.g. C_2H_4 is used in first step

(ii) ΔH shown **and** products below reactants ✓

E_a shown ✓

E_c shown $< E_a$ ✓



NOT double headed arrows but apply *ecf* for more than one double headed arrow

ALLOW one mark if two correctly labelled curves are drawn but the arrows are not shown or are incorrectly drawn

The arrows must be positioned as closely as possible to the maximum height of the curves but allow some degree of *bod*

3

[6]

11. (i) bond breaking is endothermic/
energy has to be put in to break a bond (1)

1

(ii) bonds broken: $3(\text{C-H}) + (\text{C-O}) + (\text{O-H}) + 1.5 (\text{O=O}) = 2781 \text{ kJ}$ (1)
bonds made: $2(\text{C=O}) + 4(\text{O-H}) = 3470 \text{ kJ}$ (1)

$\Delta H_c = -689 \text{ (kJ mol}^{-1}\text{)}$ (1)

3

[4]

- | | | | | | |
|-----|-----|---|--|---|-----|
| 12. | (a) | (i) | (heat/energy change) when 1 mole of substance is formed (1)
from its elements (1) | 2 | |
| | | (ii) | 1 atm/101 kPa and a stated temperature/25 °C/298 K (1) | 1 | |
| | | (iii) | $C(s) + \frac{1}{2} O_2(g) \rightarrow CO(g)$
balanced equation forming 1 mol CO (1)
state symbols (1) | 2 | |
| | | (iv) | cycle drawn/sum of $\Delta H(\text{products}) - \Delta H(\text{reactants})$ (1)
$-75 - 242 + x = -110$ (1)
$\Delta H = (+)207 \text{ kJ mol}^{-1}$ (1) | 3 | |
| | (b) | production of margarine/ammonia/Haber process (1) | | 1 | [9] |
-
- | | | | | |
|-----|--|--|--|-----|
| 13. | $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$
formulae ✓
balancing ✓
ignore state symbols | | | [2] |
|-----|--|--|--|-----|
-
- | | | | | |
|-----|---|--|--|-----|
| 14. | (enthalpy/ energy/ heat change) when 1 mole of substance/
element/ compound ✓ (NOT absorbed)
is completely burnt/ burnt in excess oxygen ✓
under standard conditions (if conditions stated they must be
correct) ✓ | | | [3] |
|-----|---|--|--|-----|
-
- | | | | | |
|-----|-------|---|---|-----|
| 15. | (i) | (enthalpy change) when 1 mole of compound is formed ✓
from the constituent elements ✓ | 2 | |
| | (ii) | $6C(s) + 7H_2(g) \rightarrow C_6H_{14}(l)$
correct formulae and balancing ✓
state symbols ✓ | 2 | |
| | (iii) | temperature 25°C/ 298K/ a stated temperature (if justified)
pressure 1 atm/ 100 kPa/ 101 kPa ✓ | 1 | [5] |

16. diagram to show
 lines to show energy level at start above that at end of reaction ✓
 ΔH labelled between reactants and products ✓
 E_a labelled from reactants to top of energy 'hump' ✓
 [3]
17. correct Hess' cycle ✓
 $x - 890 = -572 - 394$ ✓
 $x = -76 \text{ (kJ mol}^{-1}\text{)}$ ✓
 [3]
18. (i) $1652/4 = 413 \text{ (kJ mol}^{-1}\text{)}$ ✓ 1
 (ii) $(C\Box C) + 6 (C\Box H) = 2825$ ✓
 $(C\Box C) = 2825 - 6(413) = 347 \text{ (kJ mol}^{-1}\text{)}$ ✓ 2
 [3]
19. (a) (a reaction) that releases energy/ (a reaction) that releases heat/ a reaction with a negative ΔH (1) 1
- (b) (i) diagram to show
 upward **hump** (1)
 $CO_2 + (2)H_2O$ / carbon dioxide and water below
 reactants (1) 2
- (ii) E_a marked (1)
 if an arrowhead is included, it must be upwards 1
 [4]
20. (a) (heat/ energy change) when 1 mole of substance is
 formed (1)
 from its elements (1) 2
- (b) $C(s) + 2H_2(g) \rightarrow CH_4(g)$
 balanced equation (1)
 state symbols (1) 2

- (c) cycle drawn/ sum of enthalpy changes products – sum of enthalpy changes reactants (1)
 $-75 - 242 + x = -110$ (1)
 $\Delta H = 207 \text{ (kJ mol}^{-1}\text{)}$ (1) 3
- (d) any industrial use, examples include
 manufacture of ammonia/ for Haber process 1
 manufacture of margarine/ hydrogenation of alkenes

[8]

21. (i) to break a bond energy has to be put in/ ✓
 breaking bonds is endothermic 1
- (ii) energy needed to break 1 mole of bonds ✓
 in the **substance** in the gaseous state ✓ 2
- (iii) bonds broken:
 $3(\text{C-H}) + (\text{C-O}) + (\text{O-H}) + 1\frac{1}{2} (\text{O=O}) = 2781 \text{ kJ}$ ✓
 bonds made:
 $2(\text{C=O}) + 4(\text{O-H}) = 3470 \text{ kJ}$ ✓
 $\Delta H_c = -689 \text{ (kJ mol}^{-1}\text{)}$ ✓ 3
- (iv) actual bond enthalpies may be different from average values ✓
 conditions are not standard / methanol/ water is a liquid under standard conditions ✓ 2

[8]

22. (i) (enthalpy/ energy change) when 1 mole of substance/compound formed ✓
 from its elements ✓
 under standard conditions ✓ (if conditions quoted must be correct – 25 °C/298 K, 1 atm/100 kPa/101 kPa) 3
- (ii) $\text{Mg(s)} + \text{N}_2(\text{g}) + 3\text{O}_2(\text{g}) \rightarrow \text{Mg(NO}_3)_2(\text{s})$ ✓
 balanced species ✓
 state symbols ✓ 2
- (iii) cycle ✓
 $x - 791 = -602 - 2(33)$ ✓
 $x = 123$ ✓ 3

[8]

23. (i) reaction carried out at 298K and 1 atm pressure (or other relevant units) (1) 1
- (ii) enthalpy change when 1 mole (1)
(of substance) is burnt in excess oxygen (1) 2
- (iii) $4\text{CO}_2 + 5\text{H}_2\text{O}$ at lower energy than reagents (1)
 E_a marked correctly (1)
 ΔH marked correctly (1) 3
- [6]**
24. (i) $4\text{C(s)} + 5\text{H}_2\text{(g)} \rightarrow \text{C}_4\text{H}_{10}\text{(g)}$
reagents and products (1)
state symbols (1) 2
- (ii) $4\text{C} + 5\text{H}_2 \xrightarrow{X} \text{C}_4\text{H}_{10}$
 $4(-394) \quad 5(-286) \quad -2877$
 $4\text{CO}_2 \quad 5\text{H}_2\text{O}$
cycle (1)
correct values (1)
answer (1)
 $X - 2877 = 4(-394) + 5(-286)$
 $X = -129 \text{ (kJ mol}^{-1}\text{)}$ 3
- [5]**
25. (a) (i) bonds broken
 $(\text{N} - \text{N}) + (\text{O} = \text{O}) + (\text{N} - \text{H}) = 163 + 497 + 4(390) = 2220 \text{ (kJ mol}^{-1}\text{)} (1)$
bonds made
 $(\text{N} \equiv \text{N}) + 4(\text{OH}) = 945 + 4(463) = 2797 \text{ (KJ mol}^{-1}\text{)} (1)$
broken ΔH is +ve and made ΔH is -ve (1)
enthalpy of reaction $= -577 \text{ (KJ mol}^{-1}\text{)} (1)$ 4
- (ii) $\frac{577}{32} = 18.0 \text{ (KJ)} (1)$ 1

- (b) N-N bond is weak/ higher E_a for ammonia/ rate too slow for ammonia/
too much energy to break bonds in ammonia / hydrazine is liquid/ do
not need pressurised containers/ more moles/ lots of gas produced
by hydrazine/ more energy per mole produced by hydrazine (1) 1

[6]

26. (a) (enthalpy change) when 1 mole of substance/ element/ compound (1)
NOT energy needed
is completely burnt (1) 2

- (b) $C_3H_7OH(l) + 4\frac{1}{2} O_2(g) \rightarrow 3CO_2(g) + 4H_2O(l)$
correctly balanced equation (1)
state symbols (species must be correct) (1) 2

- (c) (i) $\Delta H = mc\Delta T$ (1)
 $\Delta H = 50 \times 4.18 \times 12.8 = 2675 \text{ (J)} = 2.68 \text{ (kJ)}$ (1) 2
ignore sign

- (ii) Mr propan-1-ol = 60 (1)
number moles = 0.00167 (1) 2

- (iii) $\Delta H = \tilde{1608} (KJ mol^{-1})$ (1) 1

- (iv) heat losses (1)
thermal capacity of beaker ignored (1)
conditions were non-standard (1)
combustion could be incomplete (1)
propan-1-ol evaporates (1)
water evaporates (1) 2

[11]

27. (i) the enthalpy change when 1 mole of compound/species/substance is formed ✓
[mention of 1 mole of *elements* negates this mark]
from its elements [NOT atoms/ions] (under standard conditions) ✓ 2

- (ii) 25°C/298K and 1 atm/ $1 \times 10^5 \text{ Pa}$ ✓ 1

[3]

28. $\text{Pb(s)} + \frac{1}{2} \text{O}_2\text{(g)} \rightarrow \text{PbO(s)}$ (balancing for 1 mol of PbO) ✓
 (state symbols) ✓u/c

[2]

29. (i) $\Delta H_f^\ominus = -718 - 3(-217)$
 $= -67 \text{ (kJ mol}^{-1}\text{)}$ (use of correct data & multiplier) ✓)
 (correct signs) ✓)
 (correct calculation of value) ✓) 3
- some possible ecf values:
- | | |
|-------|---|
| +67 | 2 |
| -501 | 2 |
| +501 | 1 |
| -1369 | 2 |
| +1369 | 1 |
- (ii) $\Delta H_f^\ominus = -718 + 10 + 2(217)$
 $= -274 \text{ (kJ mol}^{-1}\text{)}$ (use of correct data & multiplie) ✓)
 (correct signs) ✓)
 (correct calculation of value) ✓) 3
- some possible ecf values:
- | | |
|-------|-----|
| -57 | [2] |
| -284 | [2] |
| -294 | [2] |
| +424 | [1] |
| +444 | [2] |
| -491 | [2] |
| -511 | [1] |
| -708 | [1] |
| -1142 | [2] |
- for others, work through the calc: -[1] for each error.

[6]

30. $\text{I-I(g)} \rightarrow 2\text{I(g)}$ (state symbols ✓)
 (1 mole I_2 ✓)

[2]

31. No mark scheme available

32. No mark scheme available

33. No mark scheme available